

Unit 2 Atomic Structure

Isotopes Notes

Isotopes – Atoms of the same element that have _____
_____ because they have different numbers of
_____.

Example: ${}_6\text{C}^{12}$ and ${}_6\text{C}^{14}$ are isotopes of each other.

***Another way to denote isotopes: C-12 and C-14

Note: Isotopes of the same element _____.
_____. The atomic mass on the periodic
table is the _____ of all the natural
occurring isotopes of that element.

Problem: Which of these is not an isotope of Carbon?



Problem: What is the average mass of neon if 80.0% is Ne-20 and 20.0% is Ne-22?

Problem: What is the average mass of chlorine if 70.0% is Cl-37 and 30.0% is Cl-35?

Problem: What is the average mass of aluminum if 50.0% is Al-25, 30.0% is Al-26, and 20.0% is Al-27?